

Sensing of Gaseous Malodors in Landfills and Waste Treatment Plants

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Abstract:

We approached the problem of sensing gaseous pollutants and malodors originated from decomposition of organic compounds via chemoresistive gas sensors. A sensor array based on screen-printed films of mixed tin and titanium oxides, mixed tungsten and tin oxides, and zinc oxide has been tested on the main gaseous components of these malodors. N-butanol was also considered as a reference gas in the olfactometric intensity scale. Under proper working conditions the films can sensitively detect such gases either in dry or in wet environments, within the range of concentrations of interest for their monitoring. The array proved to be robust under solicitation by harmful interference gases such as CO, C₆H₆ e NO₂.

Key words: gas sensor, landfill, odor, sensor array, solid solution.

Introduction

Pollutants and malodors often accompany decomposition processes, giving rise to serious health risks and/or discomfort to the human agglomerates neighboring the plants where such gases are generated. Landfills and waste treatment plants are typical cases in which several types of emission gases or vapors must be carefully taken into account. In particular, landfill emissions are characterized by a very complex composition, while plants such as incinerators are potential sources of some kinds of pollutants, deriving from waste combustion, such as carbon monoxide and nitrogen oxides.

The study of the decomposition and waste combustion gases is of interest in several applications and a preliminary investigation to monitor landfill gas emission was performed in Ref. [1-6]. However, lack of selectivity is an issue and for that special care is needed while selecting the sensing units. Among the wide variety of decomposition gases we targeted ammonia, ethyl-mercaptan and sulphydric acid, because they are tracers of decomposition with stinging odor to our nose even down to very low concentrations.

Since the odorimetric aspect of the application is highly relevant, we also took into account the

responses of all the selected sensors to n-butanol, because this latter is the reference gas the odorimetric intensity scale, useful to fix the threshold for the human perception of an odor [7]. In this work, we addressed the problem of detecting typical gaseous products of waste storage, by means of gas sensors based on thick-films of metal oxides and mixed solid solutions of them, taking into account also the possible interference of typical combustion products. In order to investigate these problems different measurements were carried out: tests to identify the best detecting temperature for each type of sensing, the response to an increasing concentration and the possible interference with gases such as CO, C₆H₆ and NO₂.

Experimental

a. Selection of the sensors and preliminary operations

A set of four metal-oxide semiconducting films has been selected for the purpose. The materials chosen are ZnO, two solid solutions of SnO₂ and TiO₂, in proportions of 30% - 70%

respectively (named ST30), and 90% - 10% respectively (named ST90), and a solid solution of WO_3 and SnO_2 in proportions of 30% - 70% respectively (named WS30). The details of powders synthesis are reported in Ref. [8-13]. Measurements were performed with the flow-through technique in a sealed test chamber. Air and target gases were supplied by certified bottles, humidity was provided by means of a bubbler filled with distilled water. The gases chosen for gaseous malodors application are ethyl-mercaptan, ammonia, hydrogen sulfide and n-butanol as a reference gas for the odorimetric intensity, while CO , C_6H_6 and NO_2 were probed as combustion interferents. Humidity was monitored through a HIH-3610 Series Honeywell sensor. We investigated the responses of the sensors under dry ($\text{RH} < 2\%$ at 35°C) and wet ($\text{RH} = 23\%$ at 35°C) conditions.

b. Response to decomposition gases

The response of the set of sensors defined as $(G_{\text{gas}} - G_{\text{air}})/G_{\text{air}}$, where G_{gas} and G_{air} are the conductance values in gas and in air, respectively, was investigated vs. working temperature in order to determine the optimal detecting condition for ethyl-mercaptan, ammonia, hydrogen sulfide. The concentrations of the target gases were chosen in order to be comparable with the Recommended Exposure Limits (REL) by NIOSH (United States National Institute for Occupational Safety and Health). Indeed, for ethyl-mercaptan we chose the REL concentration of 0.5 ppm, 10 ppm for ammonia (REL = 25 ppm), 2 ppm for hydrogen sulfide (REL = 10 ppm).

The experimental results show diversification in the responses to the three gases. This is an indication of suitability of this set of sensors for the selective detection of the three target gases. The optimal working temperature for each gas is reported in Tab. 1.

Sensor	ETHYL MERCAPTAN 0.5ppm [$^\circ\text{C}$]	AMMONIA 10ppm [$^\circ\text{C}$]	H_2S 2ppm [$^\circ\text{C}$]
ST30	500	600	550
ST90	500	550	550
WS30	400	550	400
ZnO	400	400	400

Tab. 1: The optimal working temperature T_{best} for all analysed films.

At the best temperature for each gas, we measured the response to an increasing concentration of interest, highlighted in Figs. 1-3.

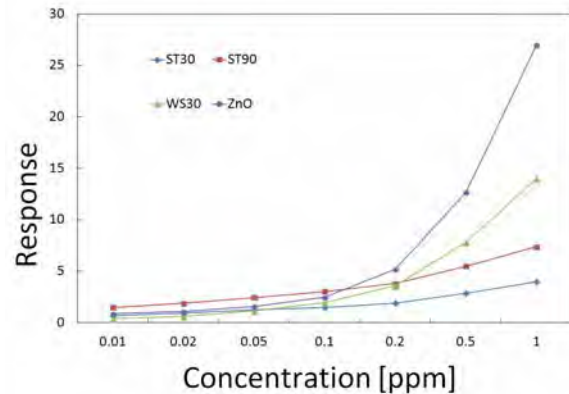


Fig. 1. Responses vs. ethyl-mercaptan within 0.01 and 1 ppm dispersed in a dry carrier.

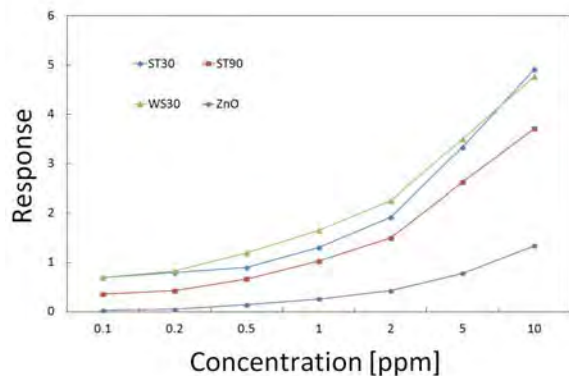


Fig. 2. Responses vs. ammonia within 0.1 and 10 ppm dispersed in a dry carrier.

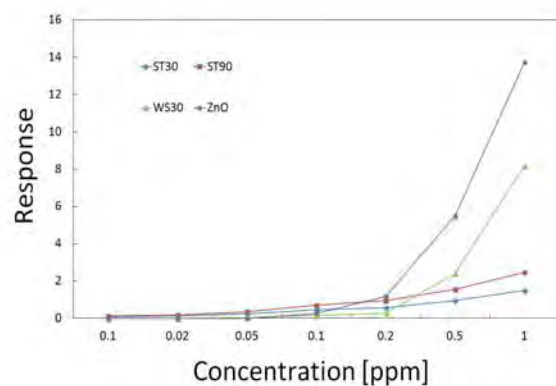


Fig. 3. Responses vs. H_2S within 0.01 and 1 ppm dispersed in a dry carrier.

Regarding ethyl-mercaptan, ammonia and hydrogen sulfide, the array encompasses sensors that are very low responsive to at least one gas, while they show significant responses to the others. This feature is a very positive indication for implementation of a high-selectivity array. In particular, ZnO at 450°C is

capable of detecting hydrogen sulfide with little interference by ethyl-mercaptan, ST30 at 600°C can selectively detect ammonia, ST90 at 500°C is capable to selectively detecting ethyl-mercaptan. WS30 is sensitive to ethyl-mercaptan and H₂S at low temperatures, while it can selectively detect ammonia at high temperature. Consequently, an array based on ZnO, ST30 and ST90 can be employed to detect the gases responsible for malodors in waste storage locations. The sensitivity of ZnO and WS30, defined as the derivative of the response vs. gas concentration, strongly increases after 0.2 ppm of H₂S. This is a useful feature since health risks becomes significant at such concentrations. For the other sensors vs. the other gases, the response is quite high to concentrations below the exposure limits.

The response of the sensors was also tested vs. the highest concentration of each gas dispersed in a wet carrier (RH = 23% at 35°C). Significant decrease in the response was recorded though it does not invalidate the capability of detection of the target gases (Fig. 4).

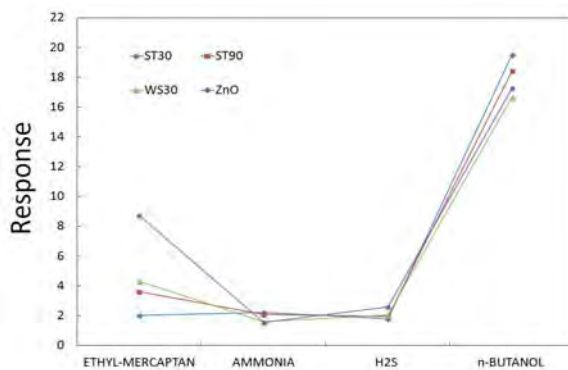


Fig. 4. Responses to 0.5 ppm of ethyl-mercaptan, 30 ppm of *n*-butanol, 10 ppm of ammonia, 5 ppm of acetaldehyde, and 1 ppm of H₂S dispersed in a wet carrier.

c. Response to *n*-butanol

The most part of molecules which are able to produce odor, are characterized by a very low olfactory threshold, then odor is well sniffed even if the gas concentration in air is negligible. This means that the odor intensity is not proportional to the gas concentration. In fact, due to the synergic and/or inhibiting effects between odorants, the relationships between the odor concentration and the chemical concentrations of the compounds are difficult to realize and reproduce. The odor concentration measurements can be carried out by three methods: analytical determination through mass spectroscopy or marker substances individuations, electronic nose and

olfactometry. This work is based on the last technique. In fact olfactometry is currently used worldwide in odor quantification because of it gives more useful information than the other two methods [14,15]. A test panel formed by selected and experienced persons, sniffs odorous gas samples conveniently diluted by odorless air on base of determined relations so that each sample is presented to the panel following a decreasing dilution series. In this way the panel identifies the “odor threshold” and it is possible to digitize a sensation in order to create an odor intensity scale (Odor Units). Relying essentially on human expertise, it is considered a time-consuming and expensive method when used frequently. In this work we needed a reference gas for the odorimetric intensity scale in order to compare the responses obtained with gaseous malodors. Therefore, we chose *n*-butanol and tested the sensors array with a concentration of 5 ppm, which corresponds to a weak odor (about 12 OU – Odor Units) [16]. In the case of *n*-butanol the temperature test shows a homogeneous behavior for all sensing films (Fig. 5) and the working temperature values obtained are reported in Table 2.

Sensor	N-BUTANOL 5ppm [°C]
ST30	550
ST90	550
WS30	600
ZnO	400

Tab. 2: The optimal working temperature *T*_{best} for *n*-butanol measurements.

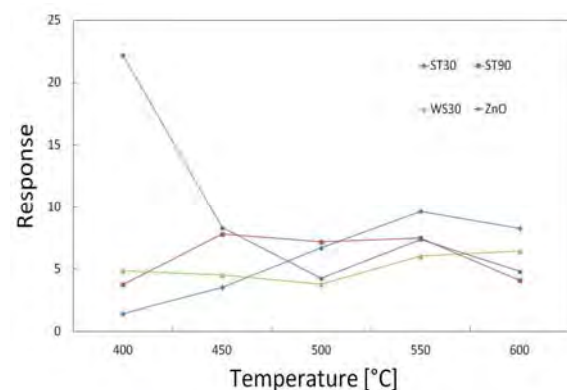


Fig. 5. Response of the set of sensors to 5 ppm of *n*-butanol vs. temperature in dry condition.

Then, once the best temperature for each type of sensor has been fixed, we measured the response to an increasing concentration of *n*-butanol, as highlighted in Fig. 6.

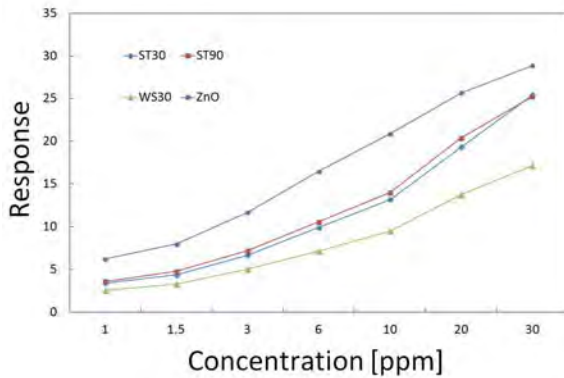


Fig. 6. Responses vs. n-butanol within 1 and 30 ppm dispersed in a dry carrier.

Since the sensors were capable of detecting concentrations of n-butanol even lower than 1 ppm, which corresponds to about 2 OU, one can conclude that sensing units are suitable not only for environmental monitoring but also for odorimetric applications. About this, since the human perception threshold is about 10 OU, it can be affirmed that the odorimetric performance of the sensors vs. n-butanol appears to be way better than the human olfact.

d. Response to harmful interfering gases

The responses of the sensors were also tested vs. CO, C₆H₆ and NO₂, since they are harmful interfering gases in zones where waste treatment plants are located. The concentrations were chosen as follows: 10 ppm for CO, 2 ppm for C₆H₆ and 4 ppm NO₂, these being concentrations high enough to simulate strong interference from combustion plants near the landfills.

The results are summarized in Figs. 7-9, where the responses of the sensors vs CO, C₆H₆ and NO₂ at several working temperatures are reported.

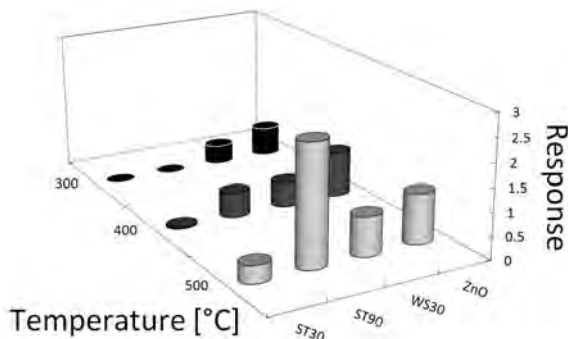


Fig. 7. Responses of the sensors to 10 ppm of CO at three different temperatures.

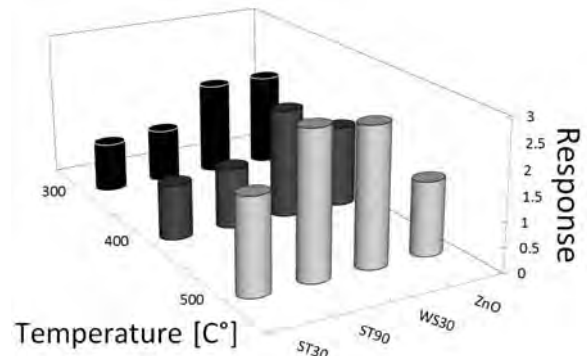


Fig. 8. Responses of the sensors to 2 ppm of C₆H₆ at three different temperatures.

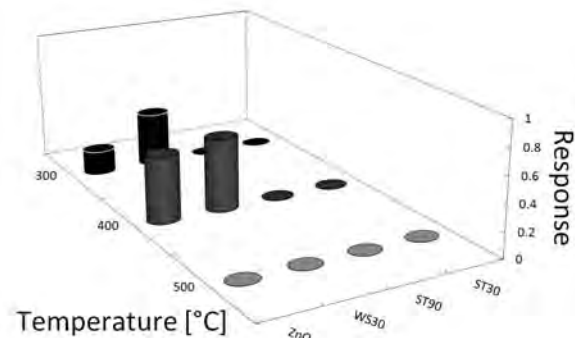


Fig. 9. Responses of the sensors to 4 ppm of NO₂ at three different temperatures.

As can be seen, all sensors show very low responses to these relatively high amounts of interfering gases. Indeed, the majority of responses are lower than the unity and in the worst case the response is less than 3, these latter being values low enough to be considered negligible in comparison with the responses to ethyl-mercaptan, ammonia and hydrogen sulfide. Therefore, the selected array of sensors can be employed in an array capable to selectively detecting the main gaseous malodors in landfills, without the interference of combustion gases that can harmfully affect the measurements.

Conclusions

The problem of sensing harmful gas pollutants and malodors typical of decomposition of organic substances has been addressed via chemoresistive gas sensors. A set of four sensing materials (ZnO, two mixed solutions of Sn and Ti oxides and a mixed solution of W and Sn oxides) has been tested in dry and wet conditions. It resulted that with a proper choice of the working temperatures the array can sensitively detect the target gases and quantify both the health risk and the odor intensity. The interference of typical gaseous species that can

be produced in waste treatment plants has been evaluated, and it resulted in a negligible response to CO, C₆H₆ and NO₂ by all the sensors.

Acknowledgments

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